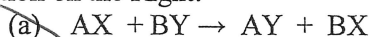
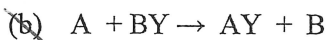
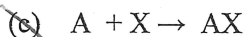
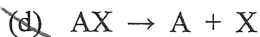


Chapter 8 Section: 2 Review

1. Match the equation type on the left to its representation on the Right.

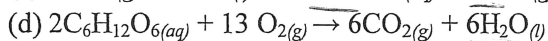
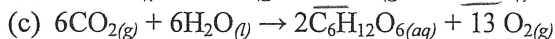
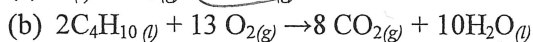
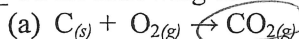
C - SynthesisD - DecompositionB - Single ReplacementA - Double Replacement2. C In the equation  $2Al_{(s)} + 3Fe(NO_3)_{2(aq)} \rightarrow 3Fe_{(s)} + 2Al(NO_3)_{3(aq)}$ , iron has been replaced by \_\_\_\_.

(a) nitrate

(b) water

(c) aluminum

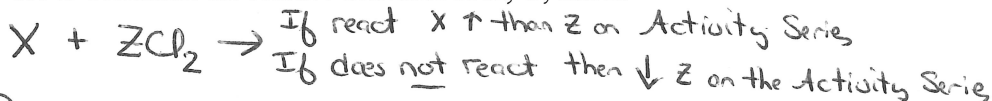
(d) nitrogen

3. A Of the following chemical equation, the only reaction that is synthesis is \_\_\_\_.

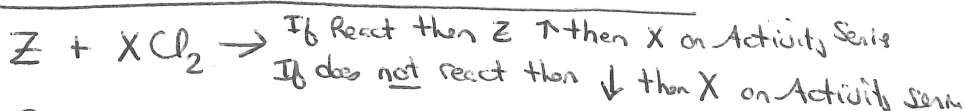
4. Identify the products when the following substances decompose:

a. Binary Compound  $\rightarrow$  individual elementsb. metal hydroxide  $\rightarrow$  metal oxide + waterc. metal carbonate  $\rightarrow$  metal oxide + CO<sub>2</sub>d. Oxyacids  $\rightarrow$  non-metal oxide + waterc. metal chlorate  $\rightarrow$  metal chloride + O<sub>2</sub>5. The complete combustion of a hydrocarbon in excess oxygen yields the products CO<sub>2</sub> & H<sub>2</sub>OChapter 8 Section: 3 Review1. List four(4) metals that will **NOT** replace hydrogen in an acid.Any metal Below H<sub>2</sub> on Activity Series  
Cu, Hg, Ag, Pt, Au

2. Consider the metals iron &amp; silver, both listed in the Activity series in RXN packet (or in text pg. 266).

Which one readily forms oxides in nature? Iron AND, which one does not? Silver.3. In the laboratory, you are given two small chips each of the unknown metals X, Y, & Z along with dropper bottles containing solutions of  $XCl_{2(aq)}$  and  $ZCl_{2(aq)}$ . Describe an experimental strategy you could use to determine the relative activities of X, Y, and Z.

Repeat w/ Y solid



Repeat w/ Y solid